described in the chapter on "Distillation Gas" can be utilised.
Gas. 0*13 cubic metre OO.2
1-54 0-19 0'67
0-01
2.71
O required. Cubic metre:
— forming
CO x 0-5=0-77 CH4x2 =0-38 II X 0-5=0-335
Cubic metre COo. 0-13 " 1-54 0-19
Cubic metre N.
O required 1*485 O deduct 0-01
O required 1*475 which involves N.....
5*55 . 2-7J
8-26 7-00
Products of combustion 1*86 + But 1'SG cubic metre carries with it only
Surplus N
This corresponds to 0'335 cubic metre 0, or
2 xO*335=:0*670 cubic metre Hcl from distillation gas.
One can arrive at the same result in another simpler and clearer manner by means of resolving the gas into its elements, which method has also been employed in the chapter on " Distillation Gas," and then considering these elements.

There exists in	Gallic metre O.	Cubic metre H.	Cubic metre 1ST.
COa 0-13 cubic metre xl	0-13	— .	—
(JO 1-54 , X-05	0-77	—	—
(Mlt 0-19	X2	—	0-38	—
IE 0-67	, X 1	—	0-67	—
0 0-01	Xl	o-oi	—	—
N 2-71	xl	—	__	2-71

*05 + 2'71+1 kilogramme 0.
5'26 cubic metre gas
consisting of . . 0*91 If now only 2*71 cubic metres N alone originates from the air, it corresponds to . . 0*72
Hence there is a surplus of 0-19 If this proceeds from tlie decomposition of steam,
there must have been
originated 2x0-19. . = 0'38
Leaving .... Q-Q7 Hd
which could only have originated from the distillation gas. It is the same figure which was obtained by the combustion method.